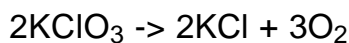


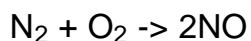
## Types of Reactions : Notes/W.S.-140

There are six types of reactions. Examples of each type (balanced) are shown below.

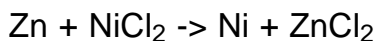
Decomposition (compound breaks apart)



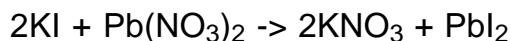
Combination (two compounds combine to form one compound)



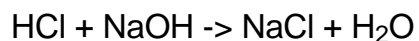
Single displacement (one element displaces another)



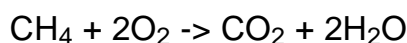
Double displacement (two elements exchange places)



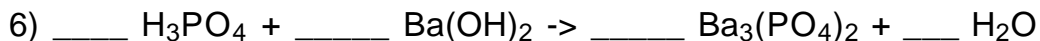
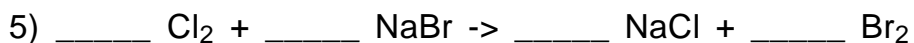
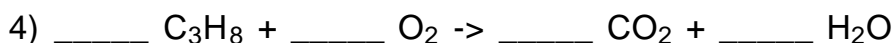
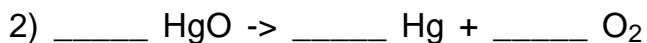
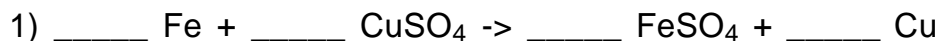
Neutralization reaction (acid plus base react to form water)

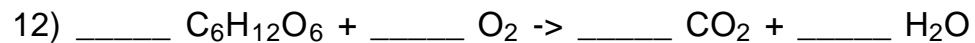
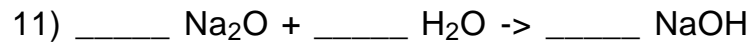
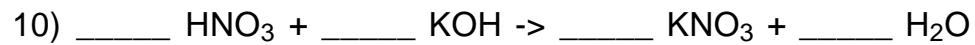
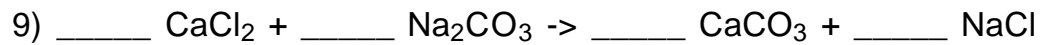
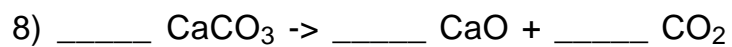
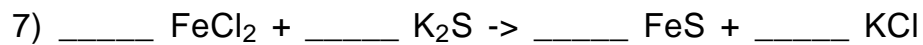


Combustion (compound burns to give off water plus carbon dioxide)



Name the type of reaction, and balance.





Answers: 1) s.d., 1,1,1,1, 2) dec, 2,2,1, 3) combination, 2,1,2, 4) combustion, 1,5,3,4, 5) s.d., 1,2,2,1, 6) neu, 2,3,1,6, 7) d.d., 1,1,1,2, 8) dec, 1,1,1, 9) d.d., 1,1,1,2, 10) neu, 1,1,1,1, 11) combination, 1,1,2, 12) combustion, 1,6,6,6.