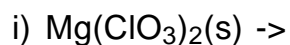
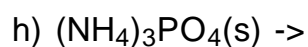
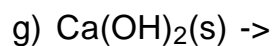
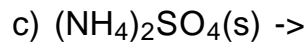
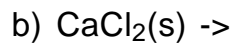


Chem12 Dissociation Equations : W.S.-20

1) Write the balanced dissociation equations.



2) Write the balanced dissociation equations for :

a) sodium chloride

b) hydrochloric acid

c) sodium hydroxide

d) ammonium acetate

e) potassium hydroxide

f) sulfuric acid

g) potassium dichromate

h) nitric acid

i) copper (II) sulfate

j) potassium permanganate

k) sodium bicarbonate

Answers : Note : All ions are in aqueous solution (aq). 1)a) $\text{Na}^+ + \text{Br}^-$, b) $\text{Ca}^{2+} + 2\text{Cl}^-$, c) $2\text{NH}_4^+ + \text{SO}_4^{2-}$, d) $2\text{K}^+ + \text{CO}_3^{2-}$, e) $2\text{Fe}^{3+} + 3\text{SO}_4^{2-}$, f) $3\text{Li}^+ + \text{PO}_4^{3-}$, g) $\text{Ca}^{2+} + 2\text{OH}^-$, h) $3\text{NH}_4^+ + \text{PO}_4^{3-}$, i) $\text{Mg}^{2+} + 2\text{ClO}_3^-$. 2)a) $\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$, b) $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$, c) $\text{NaOH} \rightarrow \text{Na}^+ + \text{OH}^-$, d) $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2 \rightarrow \text{NH}_4^+ + \text{C}_2\text{H}_3\text{O}_2^-$, e) $\text{KOH} \rightarrow \text{K}^+ + \text{OH}^-$, f) $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$, g) $\text{K}_2\text{Cr}_2\text{O}_7 \rightarrow 2\text{K}^+ + \text{Cr}_2\text{O}_7^{2-}$, h) $\text{HNO}_3 \rightarrow \text{H}^+ + \text{NO}_3^-$, i) $\text{CuSO}_4 \rightarrow \text{Cu}^{2+} + \text{SO}_4^{2-}$, j) $\text{KMnO}_4 \rightarrow \text{K}^+ + \text{MnO}_4^-$, k) $\text{NaHCO}_3 \rightarrow \text{Na}^+ + \text{HCO}_3^-$.