

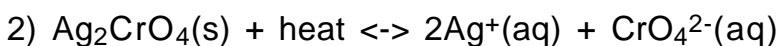
Chem12 Le Chatelier's Principle : Quiz-40

For each of the following reactions, state whether the given change will cause the equilibrium to shift toward the reactants, products or neither.



a) increase T b) decrease P c) inject more $\text{O}_2(\text{g})$

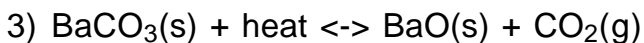
d) inject $\text{Ar}(\text{g})$ e) add a catalyst f) decrease volume



a) add more $\text{Ag}_2\text{CrO}_4(\text{s})$ b) increase T c) add $\text{AgNO}_3(\text{aq})$

d) increase P e) add $\text{NaCl}(\text{aq})$; (hint : it precipitates AgCl).

f) add $\text{Na}_2\text{CrO}_4(\text{aq})$



a) add $\text{BaCO}_3(\text{s})$ b) increase T c) increase P

d) increase volume e) add $\text{BaO}(\text{s})$ f) add $\text{CO}_2(\text{g})$

Answers : 1) r, r, p, n, n, p, 2) n, p, r, n, p, r, 3) n, p, r, p, n, r.