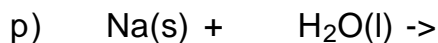
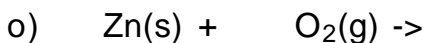
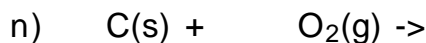
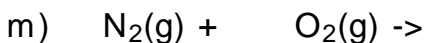
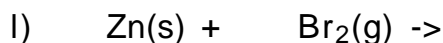
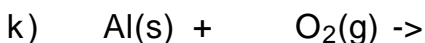
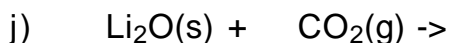
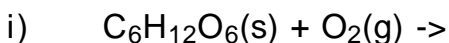
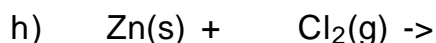
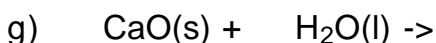
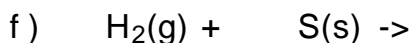
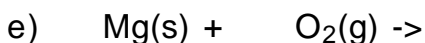
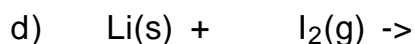
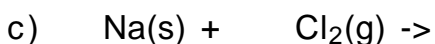
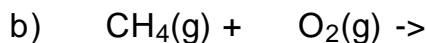
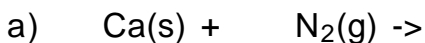


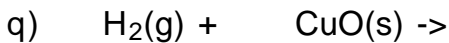
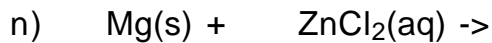
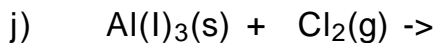
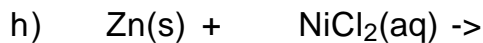
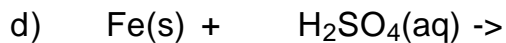
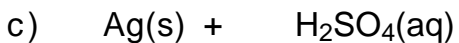
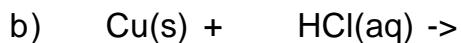
Chem11 Completing Equations 1 : W.S. - 90

(do not balance)

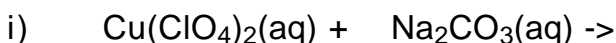
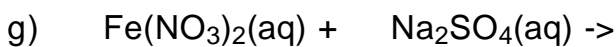
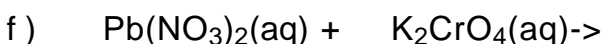
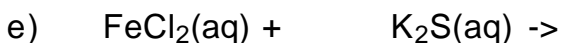
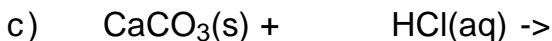
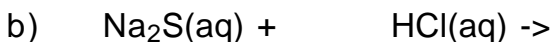
1) Complete these equations :



2) Complete these single replacement reactions. If no reaction will proceed, write N.R. (for no reaction)



3) Complete these double replacement reactions



Answers : 1)a) Ca_3N_2 , b) $\text{CO}_2 + \text{H}_2\text{O}$, c) NaCl , d) LiI , e) MgO , f) H_2S , g) $\text{Ca}(\text{OH})_2$, h) ZnCl_2 , i) $\text{CO}_2 + \text{H}_2\text{O}$, j) Li_2CO_3 , k) Al_2O_3 , l) ZnBr_2 , m) NO , $\text{NO}_2\dots$, n) CO , CO_2 , o) ZnO , p) $\text{NaOH} + \text{H}_2$, 2)a) $\text{H}_2 + \text{ZnCl}_2$, b) N.R., c) N.R., d) $\text{H}_2 + \text{FeSO}_4$, e) $\text{Br}_2 + \text{NaCl}$, f) $\text{Cu} + \text{FeCl}_2$, g) $\text{H}_2 + \text{CaBr}_2$, h) $\text{Ni} + \text{ZnCl}_2$, i) $\text{Ag} + \text{Cu}(\text{NO}_3)_2$, j) $\text{I}_2 + \text{AlCl}_3$, k) N.R., l) N.R., m) $\text{Ag} + \text{Pb}(\text{NO}_3)_2$, n) $\text{Zn} + \text{MgCl}_2$, o) N.R., p) $\text{Cu} + \text{Al}_2(\text{SO}_4)_3$, q) $\text{Cu} + \text{H}_2\text{O}$, r) N.R. 3)a) $\text{Al}_2(\text{CO}_3)_3 + \text{NaCl}$, b) $\text{H}_2\text{S} + \text{NaCl}$, c) $\text{CaCl}_2 + \text{H}_2\text{CO}_3$, d) $\text{PbSO}_4 + \text{HNO}_3$, e) $\text{KCl} + \text{FeS}$, f) $\text{KNO}_3 + \text{PbCrO}_4$, g) $\text{FeSO}_4 + \text{NaNO}_3$, h) $\text{Zn}(\text{NO}_3)_2 + \text{AgBr}$, i) $\text{CuCO}_3 + \text{NaClO}_4$, j) $\text{PbBr}_2 + \text{LiClO}_4$, k) $\text{NaCl} + \text{CaCO}_3$.