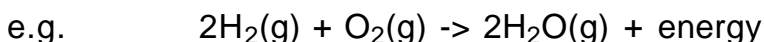


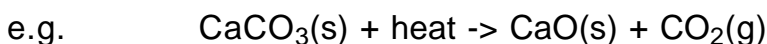
Types of Reactions : Notes/W.S.-85

There are several main types of chemical reactions.

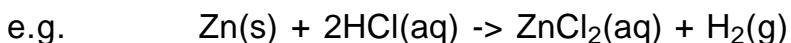
Synthesis In this type of reaction, two atoms or molecules combine to form a new molecule.



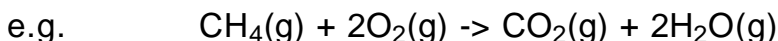
Decomposition In this type of reaction, a compound breaks apart to form two or more new compounds.



Single Replacement In this type of reaction, an element replaces another element in a compound. (the element that is higher in the activity series will replace one that is lower. see Notes-10)



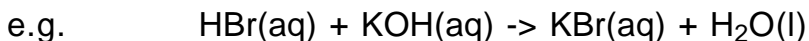
Combustion In this type of reaction, a compound combines with oxygen. A large amount of energy is usually released.



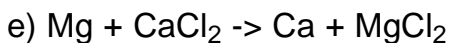
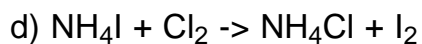
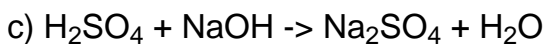
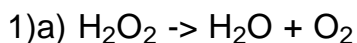
Double Replacement Two elements change places.

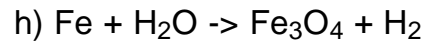
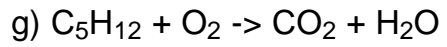


Water-forming (neutralization) An acid (contains H) and a base (contains OH) form a salt and water when mixed.



Exercise: Classify the following (unbalanced) reactions.





Answers: 1)a) decomposition, b) synthesis, c) water-forming, d) single replacement, e) reaction not possible, f) decomposition, g) combustion, h) single replacement, i) double replacement, j) water forming.