# Periodic Table : Quiz -30

1) Explain how Mendeleev discovered the periodic table.

2) Explain why atoms in a group have similar properties.

3) What change did Moseley make to Mendeleev's table?

4) What is the ionization energy?

5) Give three properties of metals.

6) Circle the correct answer.

a) Which is the most metallic element?

i) P ii) In iii) Sb

b) Which atom is the largest?

i) Ca ii) Ba iii) Br

c) Which atom has the highest ionization energy?

i) K ii) Cs iii) Fe

7) Classify the following as; solid, liquid, or gas (at 20.°C).

i) Rn ii) Br iii) As

8) Classify the following as; metal, non-metal, or metalloid.

i) Si ii) Al iii) Se

9) Classify the following as; alkali metal, alkaline earth metal, halogen, or noble gas.

i) I ii) Sr iii) Kr iv) Rb

Answers: 1) He arranged the elements in order of increasing mass and found that elements with similar properties appeared periodically., 2) The atoms of a group have the same number of valence electrons. The valence electrons determine, to a large extent, the properties of the atom., 3) Moseley arranged the elements in order of increasing atomic number., 4) This is the energy required to remove a valence electron., 5) shiny, good conductor, malleable, ductile, 6)a) ii, b) ii, c) iii, 7) gas, liquid, solid, 8) metalloid, metal, non-metal, 9) halogen, alkaline earth, noble gas, alkali metal.