

## Chem11 Intro The Periodic Table : W.S.-60

Use the Periodic Table and notes to answer the following questions :

1) Identify all eight elements that form diatomic molecules.

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2) Identify the element which is : a) tetratomic (has four atoms) \_\_\_\_

b) octatomic (has eight atoms) \_\_\_\_

3) Name the element which can act like a metal or a non-metal. \_\_\_\_

4) Rows of the periodic table are called \_\_\_\_ . Columns are called

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5)a) Name the three main classifications of elements \_\_\_\_\_

b) Which of the above classifications is the largest?

6) Select the most metallic element.

a) sulfur or selenium

b) gold or lead

c) iron or nickel

d) platinum or palladium

e) radium or strontium

f) barium or francium

7)a) The number of protons in an atom is called the

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b) The atomic mass is defined as \_\_\_\_\_

c) The atomic number for chromium is \_\_\_\_ . The atomic weight for chromium is \_\_\_\_ A.M.U.

8) What is true of elements in a group? \_\_\_\_\_

9) Name the atoms found in the third period. \_\_\_\_\_

10) What is the physical state of elements in the last column? \_\_\_\_

11) Why are the lanthanides and actinides placed at the bottom of the table?

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12) What is the state at room temperature of the following?

a) Br \_\_\_\_\_      b) Na \_\_\_\_\_      c) F \_\_\_\_

13) What is the significance of the "staircase" in the table ?

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14) What are the elements silicon, germanium, arsenic and antimony called ? \_\_\_\_\_

Answers : 1) H, N, O, F, Cl, Br, I, At, 2)a) P, b) S, 3) H, 4) Periods, Groups, 5)a) metals, non-metals, semiconductors, b) metals, 6)a) Se, b) Au, c) Fe, d) Pt, e) Ra, f) Fr, 7)a) atomic number, b) the average number of protons and neutrons in the nucleus, units are in A.M.U.s, c) 24, 52, 8) They have similar properties, 9) Na, Mg, Al, Si, P, S, Cl, Ar, 10) gas, 11) It is done so the table is not too long, 12)a) liquid, b) solid, c) gas, 13) It separates metals from non-metals, 14) semiconductors.